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Metallic Phase-Change Materials for Solar Dynamic Energy Storage Systems

R. J. Lauf
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Metals and Ceramics Division

METALLIC PHASE-CHANGE MATERIALS FOR SOLAR DYNAMIC
ENERGY STORAGE SYSTEMS

R. J. Lauf and C. Hamby, Jr.

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METALLIC PHASE-CHANGE MATERIALS FOR SOLAR DYNAMIC ENERGY STORAGE SYSTEMS*

R. J. Lauf and C. Hamby, Jr.

ABSTRACT

Solar (thermal) dynamic power systems for satellites require a heat storage system that is capable of operating the engine during eclipse. The conventional approach to this thermal storage problem is to use the latent heat of fluoride salts, which would melt during insolation and freeze during eclipse. Although candidate fluorides have large heats of fusion per unit mass, their poor thermal conductivity limits the rate at which energy can be transferred to and from the storage device. System performance is further limited by the high parasitic mass of the superalloy canisters needed to contain the salt.

This report describes a new thermal storage system in which the phase-change material (PCM) is a metal (typically germanium) contained in modular graphite canisters. These modules exhibit good thermal conductivity and low parasitic mass, and they are physically and chemically stable. Prototype modules have survived over 600 melt/freeze cycles without degradation.

Advanced concepts to further improve performance are described. These concepts include the selection of ternary eutectic alloys to provide a wider range of useful melting temperatures and the use of infiltration to control the location of liquid alloy and to compensate for differences in thermal expansion.

1. INTRODUCTION

Solar dynamic power systems¹⁻³ using either Brayton or Stirling engine cycles require a heat storage system capable of powering the engine at optimal efficiency during eclipse. Current baseline systems using the latent heat of melting of fluoride salts are

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limited by, among other things, the poor thermal conductivity of the salt and the large parasitic mass of the metal containers. For example, the heat of fusion of LiF-20% CaF₂ is 815 kJ/kg, but when the mass of the Haynes 188 canister is included, the system storage capacity becomes only 300 kJ/kg (Table 1).

Table 1. State-of-the-art heat storage system for Brayton cycle solar dynamic systems

<i>PCM</i>	
Material	LiF-20% CaF ₂
Melt temperature	1040 K
Heat of fusion	
Specific	815 kJ/kg
Volumetric	1.71 kJ/cm ³
Thermal conductivity	
Solid	4.0 W/m·K
Liquid	1.6 W/m·K
Density, liquid	2.1 g/cm ³
Void formation	22% $\Delta V/V_l$
<i>Container</i>	
Material	Haynes 188
Maximum temperature	~1100 K
<i>System</i>	
Specific heat storage ^a (maximum)	300 kJ/kg
Volumetric heat storage ^a (maximum)	1.3 kJ/cm ³

^aEstimated using the dimensions of a heat storage canister proposed for a space station system, assuming complete melting and freezing of the PCM in each cycle.

Higher thermal conductivity (k_{th}) would improve performance by allowing more of the PCM to melt in a given time with a given ΔT . One obvious way to increase k_{th} is to use a metal as the PCM. Most liquid metals with melting temperatures in the range of interest (1100–1200 K) are extremely corrosive to other metals, suggesting the need for a ceramic container. Because ceramics are inherently brittle, the design must accommodate differential thermal expansion as well as any volume change on melting. The system must survive melting and freezing every 1.5 h for 10 years (about 58,000 melt/freeze cycles) and tolerate an occasional cooldown to ambient.

Previously, Wichner, Schmidt, and DeVan at Oak Ridge National Laboratory examined many alloy systems for possible candidate PCMs. Three types of alloys were considered:

1. congruently melting line compounds or elements, in which a single solid phase of fixed composition is in equilibrium with a single liquid phase at its melting point;
2. eutectics, in which two solid phases are in equilibrium with a single liquid phase at the eutectic temperature; and
3. solid solutions, in which a single solid phase melts incongruently between the solidus and liquidus temperatures.

Congruently melting line compounds, elements, and eutectics are particularly appropriate as PCMs, because they theoretically release energy isothermally on solidification. Solid solutions are less appropriate as PCMs because of macrosegregation, which is likely to occur from multiple freeze/melt cycles. This segregation might shift the melting range of the PCM.

Wichner et al. organized the candidate PCMs into four groups based on their likely chemical properties (Table 2). Of these, we selected three materials for near-term development: Ge (element), Mg_2Si -Si (eutectic), and NiSi-NiSi₂ (eutectic). The properties of these three PCMs, and those of LiF-CaF₂ for comparison, are listed in Table 3. Pure Ge melts at 1210 K, and on melting there is a volume decrease of nearly 5%. Any capsule design must therefore take into account the corresponding expansion on freezing. The Mg_2Si -Si eutectic melts at 1219 K and also appears to expand somewhat on freezing. The NiSi-NiSi₂ eutectic melts at 1239 K with little (if any) volume change.

As originally conceived these alloys were to be formed into small beads and coated with thin layers of SiC by chemical vapor deposition. The coated PCM particles would then be dispersed in a graphitic matrix that could be extruded into convenient shapes (e.g., cylinders). This technology is well known, having been developed originally for the High-Temperature Gas-Cooled Reactor⁴ (HTGR).

Although the microsphere concept could, in principle, contain the PCM reliably, its major drawback is the large parasitic mass (and volume) of coating and matrix. Therefore, we adopted a simpler design in which a large mass of PCM was contained directly in a ceramic canister. Graphite was chosen as the container material because of its excellent

Table 2. Four groups of Si alloy
phase-change materials

		T_m (K)	ΔH_m (kJ/kg)
<i>Group VIIIA - Si Systems</i>			
PdSi-Si	Eutectic	1143	480
PdSi	Compound	1363	520
NiSi-NiSi ₂	Eutectic	1239	640
NiSi	Compound	1265	750
Fe ₂ Si ₅ -Si	Eutectic	1480	1000
FeSi-Fe ₂ Si ₅	Eutectic	1485	900
Fe ₂ Si ₅	Compound	1493	1180
<i>Group VIIA - Si Systems</i>			
Mn ₁₁ Si ₁₉	Compound	1418	1130
ReSi ₂ -Si	Eutectic	1398	948
<i>Group IIA - Si Systems</i>			
Be-Si	Eutectic	1363	1350
Mg ₂ Si-Si	Eutectic	1219	800-1000
Mg ₂ Si	Compound	1358	1100
CaSi ₂ -Si	Eutectic	1253	1100
SrSi ₂ -Si	Eutectic	1273	870
SrSi-SrSi ₂	Eutectic	1317	660
SrSi ₂	Compound	1423	850
BaSi ₂	Compound	1453	630
<i>Ge-Si Solution Systems</i>			
Ge	Element	1211	510
Ge-0.2Si	Solution	1270-1400	650
Ge-0.4Si	Solution	1340-1510	830
Ge-0.6Si	Solution	1430-1580	1130
Ge-0.8Si	Solution	1560-1640	1330
Si	Element	1687	1800

Table 3. Properties of candidate phase-change materials

Alloy	T_m (K)	ΔH_m (kJ/g)	ΔV_m (%)	Comments
Ge	1210	481	-4.75	Pure element
Mg ₂ Si-Si	1219	800–1000		Eutectic
NiSi-NiSi ₂	1239	640		Eutectic
LiF-20% CaF ₂	1040	815	+22	Salt system for comparison

high-temperature properties, good thermal conductivity relative to other ceramics, and relatively low density. Graphite can be easily machined into large components that would be prohibitively difficult to make from structural ceramics such as silicon carbide or silicon nitride. The major shortcoming of graphite is helium permeability at high temperatures. Thus, containment of the working fluid (a mixture of He and Xe) must be provided by refractory metal tubes thermally coupled to the graphite.

The various options for sealing the graphite modules, joining them together, and mating the assembly to the gas manifold are presently under consideration and will be reported elsewhere.

2. FABRICATION AND TESTING OF THERMAL STORAGE MODULES

The graphite capsules used in this study were machined from two billets of Stackpole 2020 graphite* and one billet of POCO AXF-5Q.† The Stackpole 2020 was chosen because it represents a good grade of commercial graphite and had been well characterized in previous work. This graphite was available in two levels of purity, allowing us to test the effect of metallic impurities on wetting or adhesion. The POCO AXF-5Q represents a material with much finer porosity than the Stackpole 2020. This material was used after prototype capsules made from Stackpole 2020 began to lose some Ge, apparently by seepage through interconnected porosity. Properties of each type of graphite are summarized in Table 4.

*Stackpole Carbon Co., St. Marys, Penn.

†POCO Graphite, Inc., Decatur, Tex.

Table 4. Graphite properties and chemistry

Material ^a	Stackpole 2020 Grade 1	Stackpole 2020 Grade 2	POCO AXF-5Q
Density, g/cm ³	1.79	1.78	1.79
Flexure strength, MPa (psi)	34.5 (5000)	34.5 (5000)	89.7 (13,000)
Coefficient of thermal expansion, 10 ⁻⁶ /K	3.5	3.5	8.4
Pore size, μm	1.5	1.5	0.8
Grain size, μm	40	40	4
Impurities ^b			
Ash, wt %	0.11	0.11	0.10
Fe, ppm	11	110	5-30
Si, ppm	70	240	5-50
Ca, ppm	260	330	5-100
Mg, ppm	0.5	0.07	1-5
Al, ppm	16	115	10-50
V, ppm	15	<i>c</i>	100-500

^aGrade 1 and Grade 2 are internal sample numbers and are not the manufacturer's designation.

^bImpurity levels in Stackpole 2020 represent actual measurements from nearby billet sections. Values for POCO AXF-5Q are maximum limits specified by manufacturer.

^cNot determined.

Graphite capsules were made in two sizes, which were designated "test" and "prototype" geometries. The test capsule was a cylindrical crucible with a threaded closure (Fig. 1). This capsule has a maximum capacity of 85 g Ge for a theoretical storage capacity of 38 kJ. Twelve capsules were machined from each grade of Stackpole 2020. They were used for initial screening studies to determine the most promising alloys for testing in the prototype geometry.

The prototype capsules were cylindrical [7.6 cm (3 in.) OD \times 2.5 cm (1 in.) ID \times 7.6 cm (3 in.) long], and designed to approximate heat storage modules that would slip over Nb-1Zr tubes carrying the working fluid in an actual solar dynamic power unit. These capsules are therefore directly analogous to the salt canisters of the baseline PCM system. Two prototype designs were tested. The first is a hollow annulus with a maximum capacity

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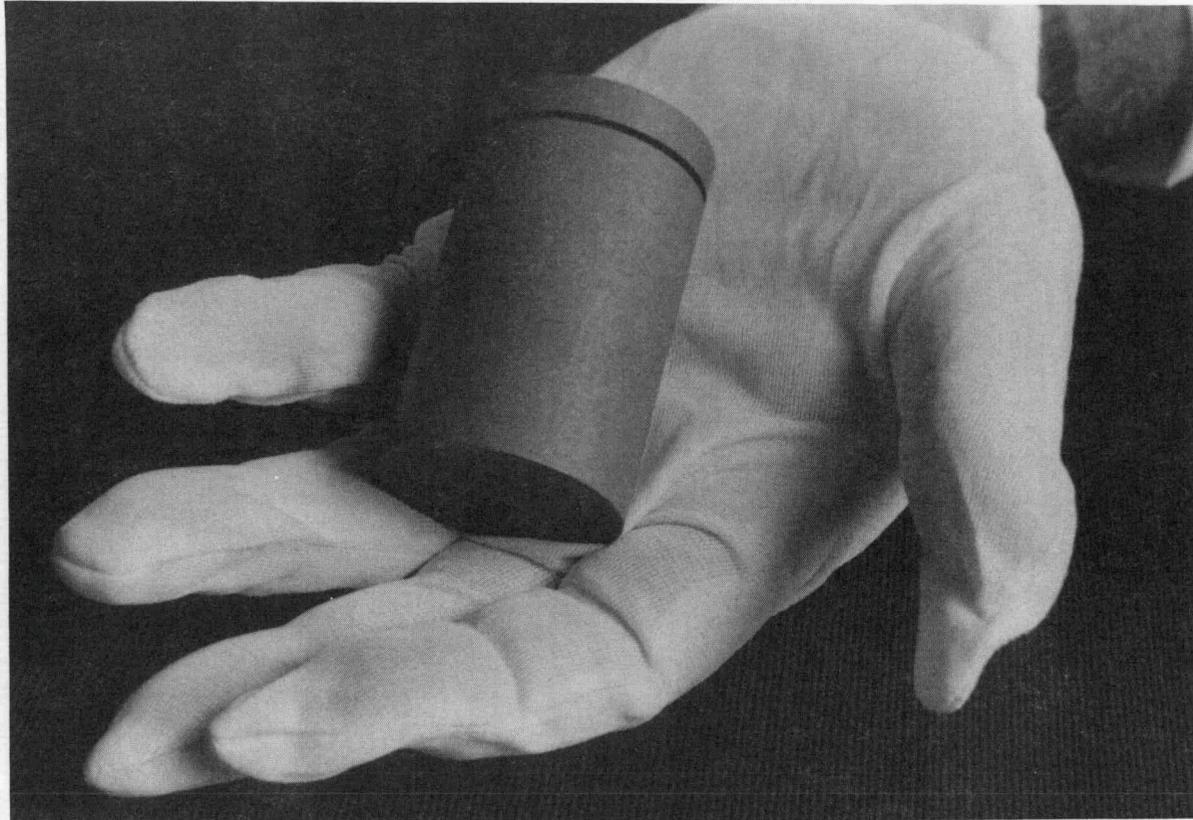


Fig. 1. Graphite test capsule with threaded closure, designed to hold 85 g Ge and store 38 kJ of latent heat at 1210 K.

of 900 g Ge and 440 kJ of latent heat. The second prototype design is a solid annulus with axially drilled cavities to contain the PCM. This module holds 750 g of Ge and stores 360 kJ of latent heat. Both capsules have threaded closures (Fig. 2).

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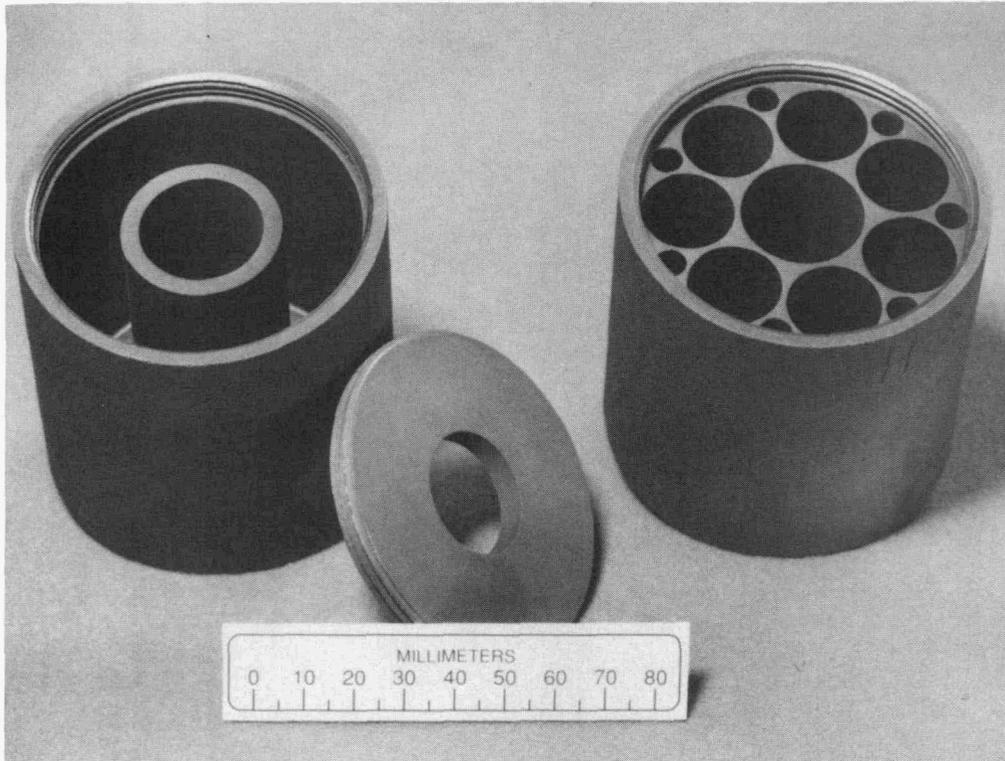


Fig. 2. Graphite prototype capsules of two designs: left, hollow annulus; right, drilled cavities.

Capsules of both test and prototype geometries were subjected to thermal cycling to test their ability not only to contain the molten PCM but also to accommodate volume changes on melting and differential thermal expansion between the PCM and the graphite capsule. Thermal cycle tests were done in a vacuum of about 10^{-6} torr. A typical cycle consisted of 1 h at 773 K, heatup at 17 K/min to 1273 K, 2 h at 1273 K, and cooldown at 11 K/min to 773 K. The furnace cycled automatically; periodically the samples were cooled to ambient, removed for inspection and weighing, and returned for additional cycling.

Initially, all specimens were arranged in the furnace cavity to melt and freeze isothermally (i.e., no forced cooling was provided, and heat transfer was predominantly radiative). Later, a stainless steel cold finger was designed to fit inside the bore of the prototype capsule and extract some of the heat with flowing nitrogen (Fig. 3). The purpose of forced cooling was to determine the influence of directional solidification on the movement of PCM within the capsule (recalling that Ge expands on freezing). In operation, layers of carbon-bonded carbon-fiber (CBCF) insulation were placed above and below the capsule in an effort to maintain a purely radial thermal gradient. Nitrogen was delivered to the cold finger through a solenoid activated by the furnace controller so that gas flowed only during the first 15 min of each cooldown cycle. This was more than adequate to extract the latent heat and completely solidify the PCM.

A complete summary of all tests (both thermal cycle and sessile drop) is given in the Appendix. In the discussion that follows, the test results are grouped according to the particular PCM to illustrate how each alloy system behaves during melting and freezing and how each interacts with its graphite container.

2.1 GERMANIUM AND ITS ALLOYS

2.1.1 Sessile Drop

The Ge-C system⁵ shows complete insolubility between the two elements. Sessile drop tests of pure Ge on porous CBCF and on a spectroscopic carbon planchette, both in vacuum at 1273 K, confirmed that pure Ge does not wet graphite at all; the sessile drop was nearly spherical and did not adhere to the carbon on cooling. Using the standard notation,⁶ the contact angle $\theta = 180^\circ$. To explore the possibility that some wetting might be desirable to improve heat transfer and to better control the location of the alloy within the container, two Ge alloys were made containing 0.5% Ti and 0.5% Si, respectively. The contact angle of Ge-0.5% Ti on a carbon planchette was 150° and that of Ge-0.5% Si was 135° , indicating that both additives increase wetting (Figs. 4 and 5).

2.1.2 Test Capsules

Pure Ge was melted in both Grade 1 and Grade 2 Stackpole graphite test capsules. The metal did not wet either grade of graphite; because of differential thermal expansion,

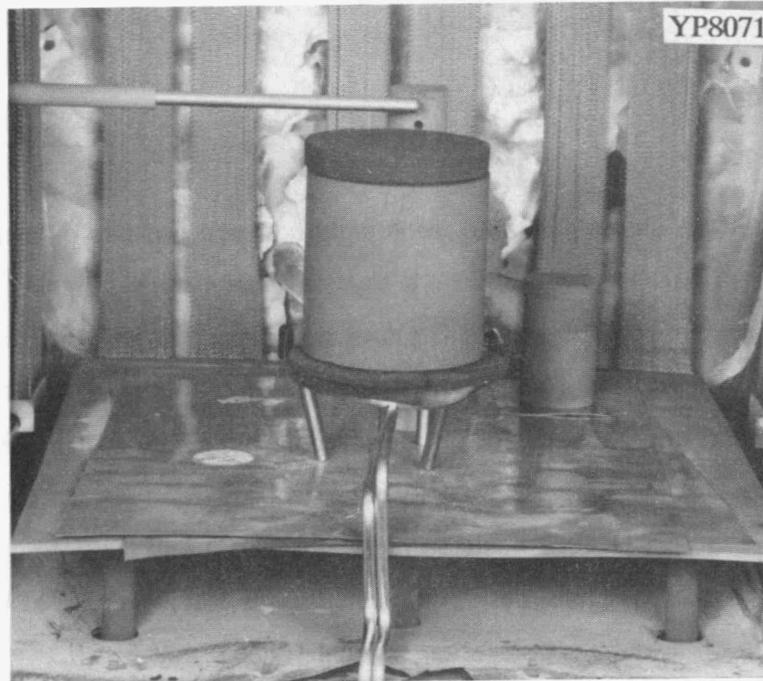
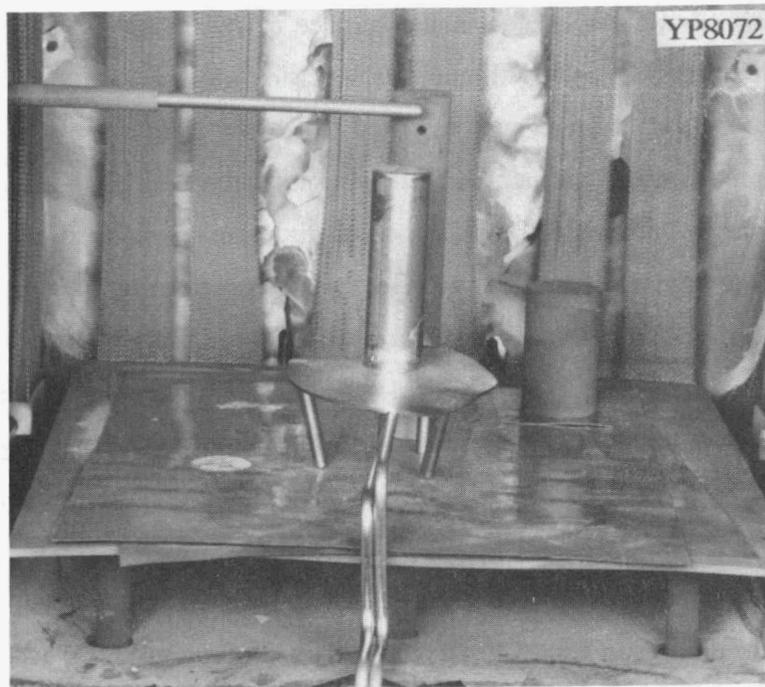


Fig. 3. Interior of vacuum furnace, showing cold finger used to cool the inside diameter of the prototype capsule during thermal cycling. Top: cold finger (note test capsule behind it on the right). Bottom: prototype capsule in place, with CBCF insulation above and below it.

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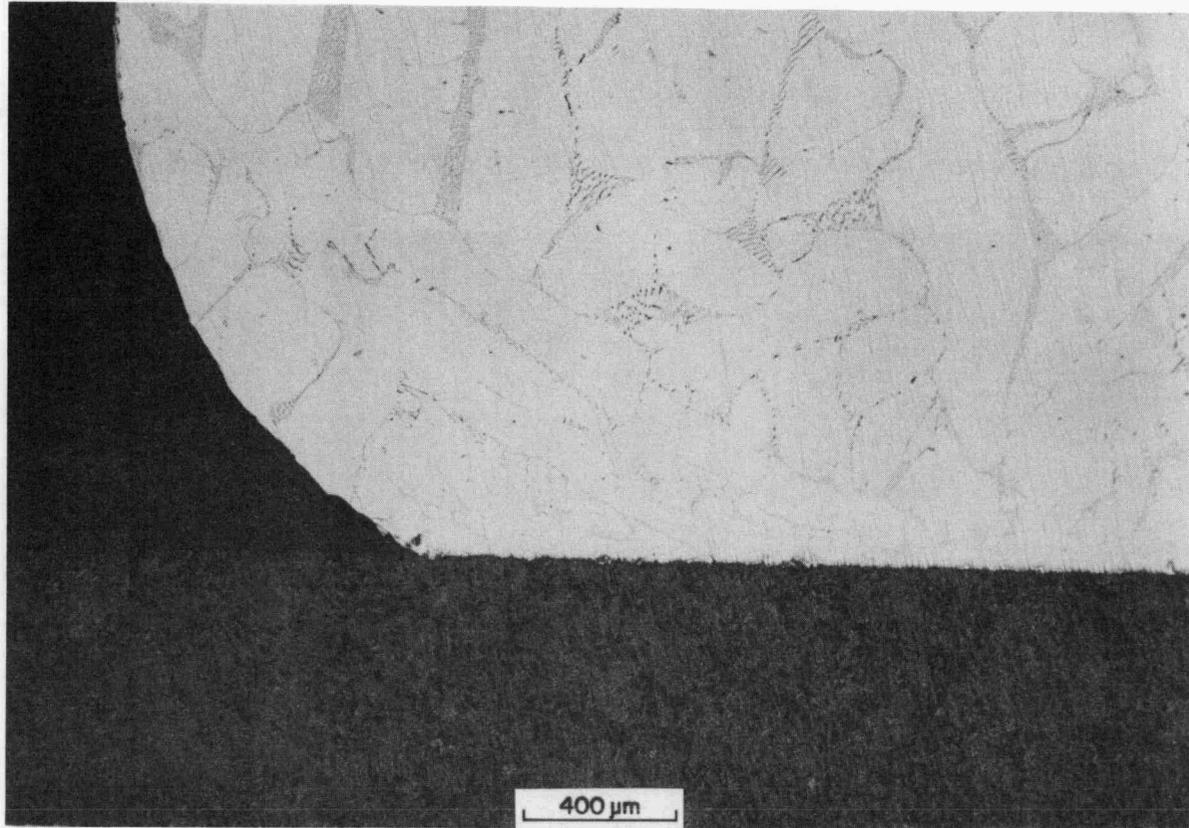


Fig. 4. Sessile drop of Ge-0.5% Ti alloy on a spectroscopic carbon planchette. Wetting angle $\theta \approx 150^\circ$.

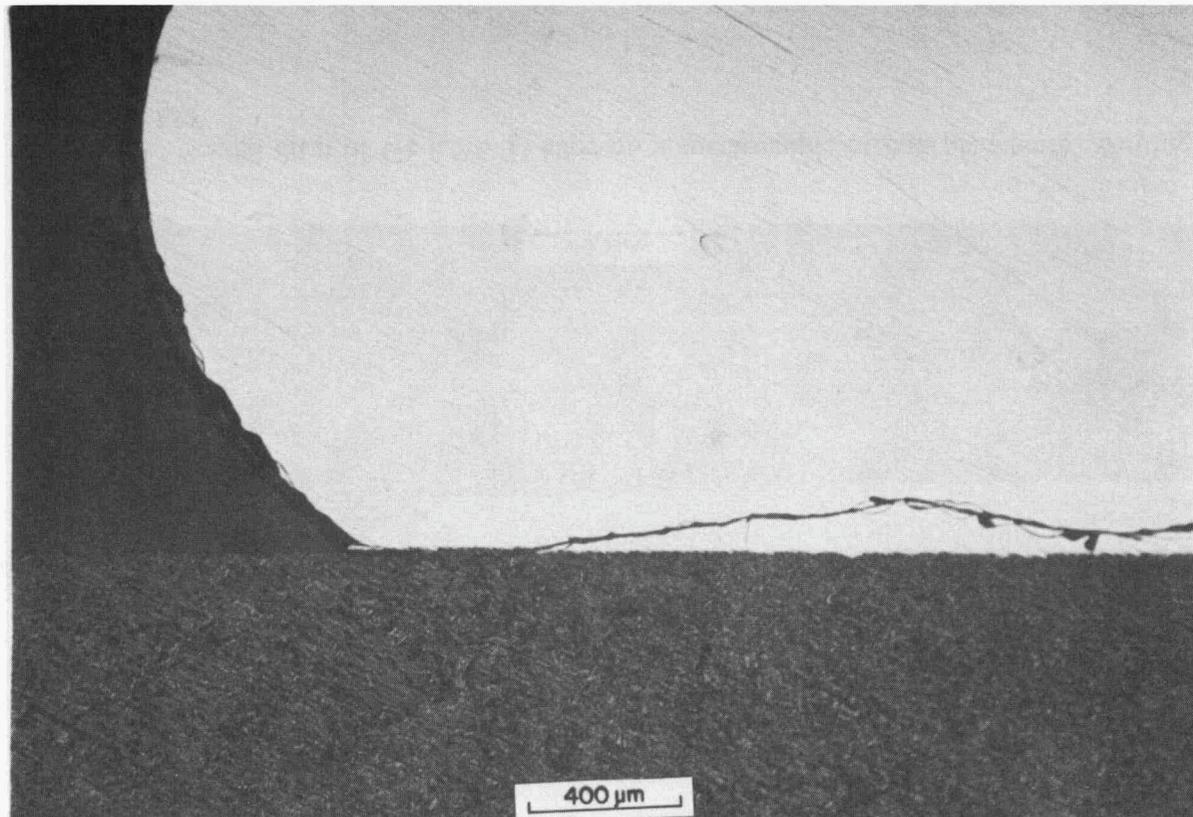


Fig. 5. Sessile drop of Ge-0.5% Si alloy on a spectroscopic carbon planchette. Wetting angle $\theta \approx 135^\circ$.

on cooling to ambient the metal was able to slide freely within the capsule. To test the adequacy of the threaded closure, the Grade 1 capsule was turned upside-down from cycles 56 through 89. No Ge leaked from the capsule, but a small amount became trapped in the innermost thread. Because of abrasion at that point during later openings, this capsule showed slightly higher weight losses than the others. This capsule is still intact after 544 cycles. The Grade 2 capsule cracked around the bottom after 544 cycles, but no Ge escaped.

An interesting feature of pure Ge is expansion on freezing, which leads to an upwelling of material at the top of the capsule (under normal gravity). This structure (Fig. 6) is nothing more than the inverse of the shrinkage cavity one normally observes at the top of a solidified metal ingot, but it serves to illustrate that expansion occurs on freezing, and any engineering design must accommodate it. Furthermore, one must recall that in a microgravity environment it becomes difficult to predict the locations of solid and liquid; yet for the capsule to survive, the liquid must always be free to communicate with the void space.



Fig. 6. Test capsule containing pure Ge, showing structure that forms by expansion of the Ge upon solidification.

The Ge-0.5% Ti alloy in a Grade 1 capsule showed some wetting during the early cycles. Later, wetting decreased, and the alloy developed a granular surface. We attribute this to loss of Ti from the melt through formation of TiC particles. A Grade 2 capsule with somewhat more PCM of this composition cracked after 52 cycles.

The Ge-0.5% Si alloy in a Grade 1 capsule showed some wetting of the graphite, which increased steadily with further cycling. After 544 cycles, the meniscus was nearly 5 mm above the bulk of the melt. A Grade 2 capsule that contained a larger quantity of this alloy cracked after 52 cycles. This failure, like the previous two, appeared to originate at the inside corner in the bottom of the capsule. This suggests that a larger radius could help decrease the stress concentration and possibly prevent these failures. It is also interesting to note that all three failures were Grade 2 material.

Although difficult to quantify (given the irregular shapes involved), it appears that neither of the Ge alloys expand as much on freezing as does pure Ge. This observation might be significant, because volume changes on freezing, if constrained, give rise to essentially irresistible forces that must be avoided in service.

An empty Grade 1 capsule was tested with the others to see how much weight loss can be attributed to graphite oxidation (due to residual gases) as opposed to PCM evaporation or leakage. Representative weight loss data (Fig. 7) show that almost all of the observed loss was accounted for by graphite oxidation. Note that absolute weight losses were used rather than relative (percent) losses, because they better illustrate this fact. Under these conditions, a capsule containing more metal would appear to show a lower percentage weight loss, obscuring the underlying phenomenon.

2.1.3 Prototype Capsules

A prototype capsule of the "hollow annulus" geometry made from Stackpole 2020 (Grade 1) that contained about 750 g Ge was tested for 482 cycles. The capsule remained intact but lost some Ge through seepage along the inside diameter, leading to large weight losses (Fig. 8). A similar capsule made from POCO AXF-5Q showed no Ge leakage and much lower weight loss than even the small Stackpole test capsules, indicating a very low rate of graphite oxidation for the AXF-5Q. The POCO capsule was tested for 358 cycles without forced cooling and an additional 44 cycles with the cold finger in place. These tests are continuing. Note that both of these capsules have two radial CBCF shims

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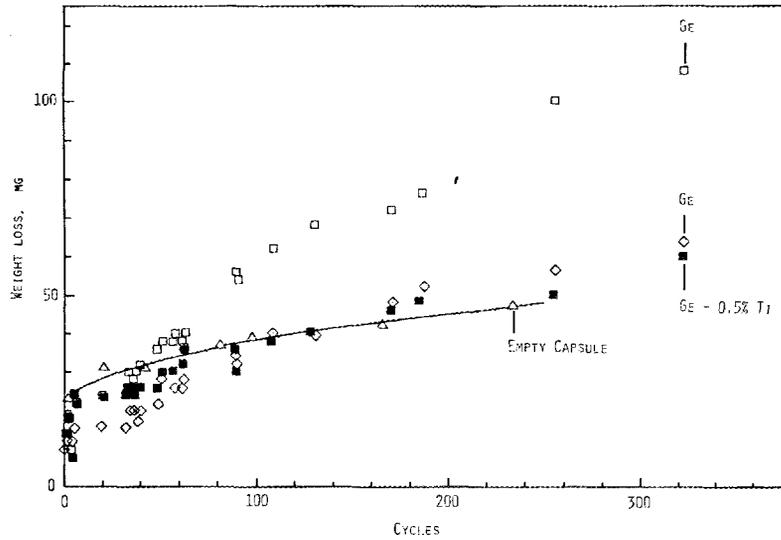


Fig. 7. Weight loss data for several test capsules. Each cycle consists of 1 h at 773 K and 2 h at 1273 K. The weight loss from an empty graphite capsule (curve with open triangles) represents the weight loss caused by graphite oxidation. Anomalous losses from one Ge sample were the result of metal adhesion to the threaded closure.

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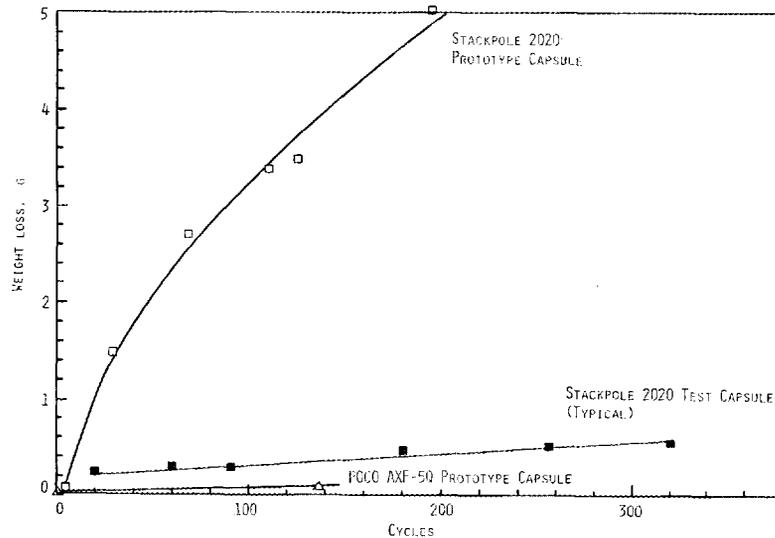


Fig. 8. Weight loss data for two prototype capsules containing Ge and for one empty test capsule. The weight loss from the 2020 prototype capsule was mainly the result of Ge loss by seepage. The loss from the test capsule represents graphite oxidation. Weight losses from the POCO AXF-5Q prototype capsule were extremely low.

running lengthwise in the cavity to divide the PCM into two D-shaped segments. This provides enough compliance to accommodate the different coefficients of thermal expansion (CTEs).

Prototype capsules of the "drilled cavities" geometry, both 2020 and AXF-5Q, were tested with germanium, and each capsule fractured on the inside diameter after three cycles. Examination of the metal PCM showed that it had frozen on the top and bottom, leaving trapped liquid in the center of the cavity. The expansion that occurred when this liquid froze broke through the capsule wall. It is possible that directional solidification control with the cold finger could eliminate this failure mode; however, it is not clear that the drilled geometry conveys any special benefits. (The original motivation to test this geometry involved greater compartmentalization of the PCM and possibly simpler mass production. It does, however, hold less Ge than the hollow annulus, so the parasitic mass is greater.) Because of the greater difficulty in preventing trapped liquid, this design is of limited interest for further development.

Table 5 compares the characteristics of germanium/graphite thermal storage modules (hollow annulus type) from the present work with those of the molten salt/superalloy system.¹ In addition to the improved storage capacity, the Ge/graphite assembly is easier to fabricate and much more tolerant of temperature excursions above the melting point of the PCM. No attempt has been made to optimize the design to minimize the parasitic mass of graphite, so further development work should lead to improved performance.

Table 5. Comparison of heat storage systems

Phase-change material:	LiF-20CaF ₂	Germanium
T _m (K)	1040	1210
ΔH _m (kJ/kg)	815	507
(kJ/cm ³)	1.71	2.70
K _h (W/m · K) solid	4.0	15
liquid	1.6	41
Container:		
Material	Haynes 188	Graphite
Maximum temperature (K)	1100	>2000
Total system:		
Maximum heat storage (kJ/kg)	300	370
(kJ/cm ³)	1.3	1.7

Crane and Bharadhwaj⁷ recently modeled the performance of several heat receiver concepts in solar-powered Brayton and Stirling engines. They compared LiF-22CaF₂ salt, several "enhanced" salts containing BN or pyrolytic carbon foams for better thermal conductivity, and pure Ge. Their calculations show that the high thermal conductivity of the Ge PCM "virtually eliminates all thermal variation" in the Stirling heat pipe receiver.

2.2 Mg₂Si-Si EUTECTIC

The Mg-Si system⁸ contains the compound Mg₂Si, which melts at 1375 K. On either side of this line compound are two simple eutectics. The Mg-Mg₂Si eutectic occurs at 911 K and 1.34 wt % Si. The Mg₂Si-Si eutectic occurs at about 1223 K and 58 wt % Si. The latter composition is the alloy of interest here as a PCM.

The Mg₂Si-Si eutectic composition cannot be made at reasonable pressures by melting a mixture of the pure elements. This is because of the large difference in melting points. At the melting temperature of Si (1703 K), vaporization losses of Mg would be excessive; therefore, we were limited to forming the eutectic in situ by mixing Si chips and Mg₂Si powder in the test capsules and melting at 1273 K. For this reason, the Mg₂Si-Si eutectic was not available in a form suitable for the sessile drop test.

2.2.1 Test Capsules

A Grade 2 Stackpole 2020 test capsule was filled with Mg₂Si powder and Si chips. After 2 h at 1273 K, the charge had melted thoroughly and settled in the capsule. More material was added and remelted at 1273 K. After the second melting, the lid adhered to the capsule and could not be unscrewed. After the third cycle, there was a small crack running about halfway around the capsule near the bottom. The test was concluded, and the capsule was sectioned lengthwise for metallographic examination. From the micrographs (Figs. 9 and 10), it is clear that this alloy readily wets graphite ($\theta \approx 0^\circ$). After only three cycles the meniscus had climbed to the top of the capsule. A reaction zone is about 100 μm deep, characterized by penetration of metal nodules and some restructuring of the graphite. If the apparent densification of the graphite in this region represents formation of SiC in the pore spaces, penetration could possibly be self-limiting. Otherwise, it will be necessary to protect the graphite from attack by applying a chemically vapor deposited (CVD) coating of SiC or a similar material. Before that option is

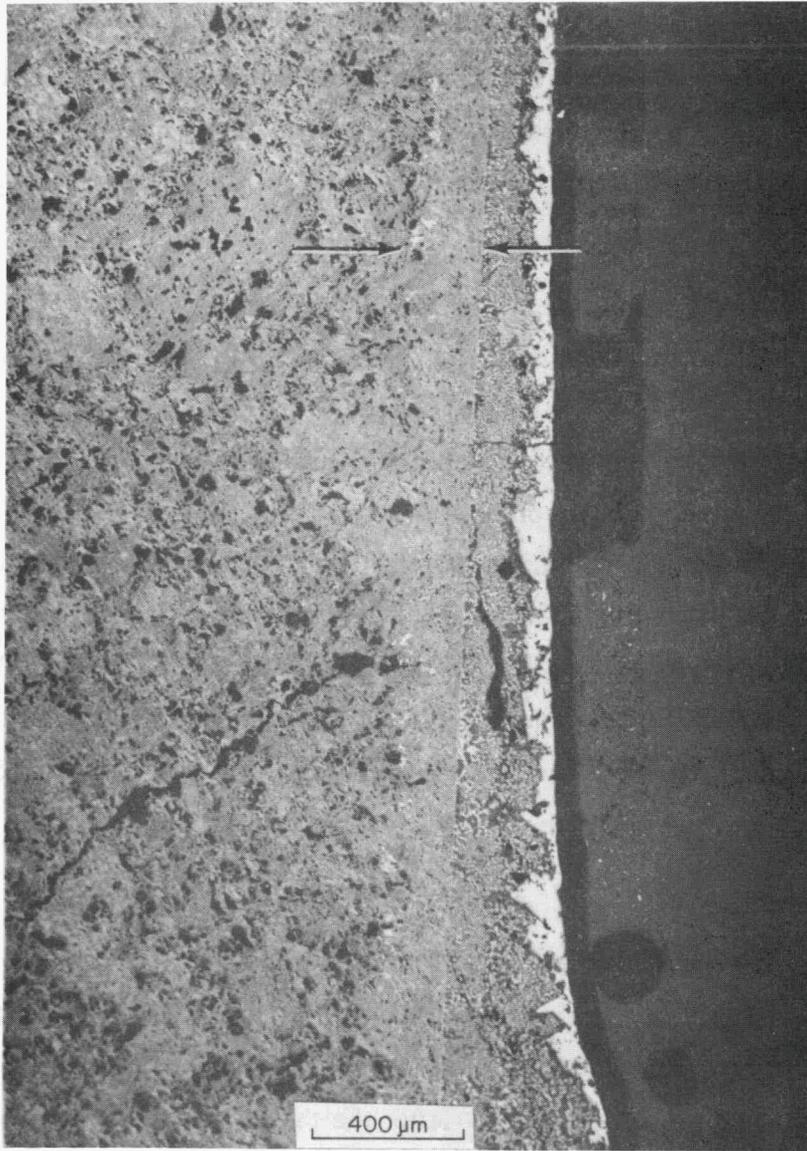


Fig. 9. Mg_2Si -Si eutectic alloy forming a thin meniscus on the inner surface of a Stackpole 2020 test capsule. Note restructured layer in the graphite (arrows).

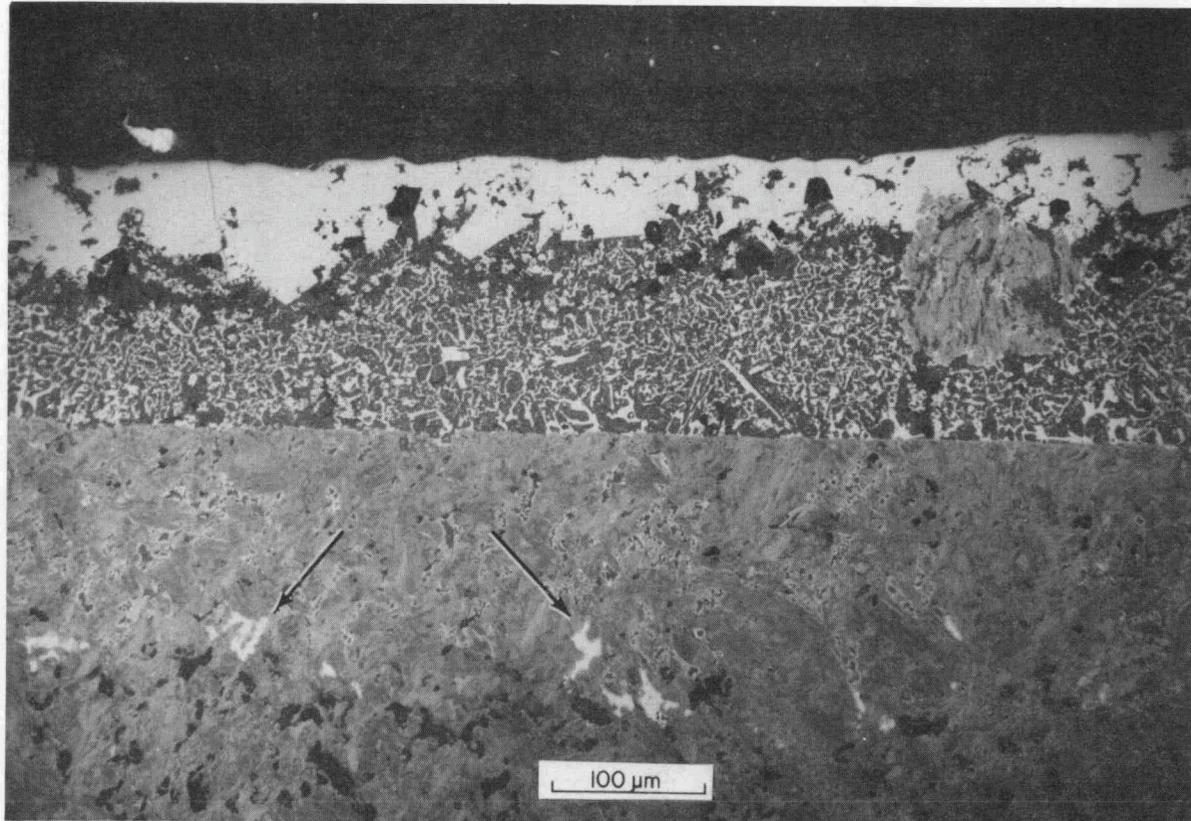


Fig. 10. Detail of Mg_2Si-Si on graphite surface. Note metal nodules about $100\ \mu m$ into the graphite (arrows).

explored, test capsules will be machined from POCO graphite to see if the attack can be eliminated by microstructural control of the graphite.

A Grade 1 Stackpole test capsule containing this alloy cracked after seven thermal cycles. Both of these early failures suggest a thermal expansion mismatch between the PCM and the graphite. Because the liquid $\text{Mg}_2\text{Si-Si}$ readily wets graphite, this PCM was used to test the infiltration approach for eliminating CTE-driven fracture. These results are discussed later.

2.3 NiSi-NiSi₂ EUTECTIC

The Ni-Si system⁹ contains a number of intermediate phases (at least seven) and numerous invariant reactions. Of interest here is the eutectic NiSi-NiSi₂, which occurs at 1238 K and 38 wt % Si. This alloy is readily made by arc-melting a mixture of the elements in a water-cooled copper hearth. Alternatively, a physical mixture of the elements will melt, with somewhat more difficulty, at temperatures as low as 1273 K in vacuum. Once melting begins, alloying generally proceeds to completion.

2.3.1 Sessile Drop

The sessile drop of arc-melted NiSi-NiSi₂ eutectic on a spectroscopic carbon planchette, held at 1273 K for 4 h in vacuum, had a contact angle of about 110°. Metallographic examination (Fig. 11) showed that some metal had penetrated into the carbon planchette. The metal bead was extensively cracked and appeared to have dissolved some carbon, which was rejected as a second phase on cooling. This is consistent with previous observations that certain low-melting nickel and cobalt alloys exhibit remarkably high carbon solubilities in the liquid phase.

2.3.2 Test Capsules

NiSi-NiSi₂ appears to have a much lower CTE than graphite. Test capsules containing this alloy invariably fractured on their first cooldown to ambient. At the same time, however, capsules appeared to survive any number of cycles between 773 and 1273 K, because there was no evidence of prior leakage from the broken capsules. (Unlike the Mg-Si capsules that simply developed small cracks, the Ni-Si capsules were catastrophically

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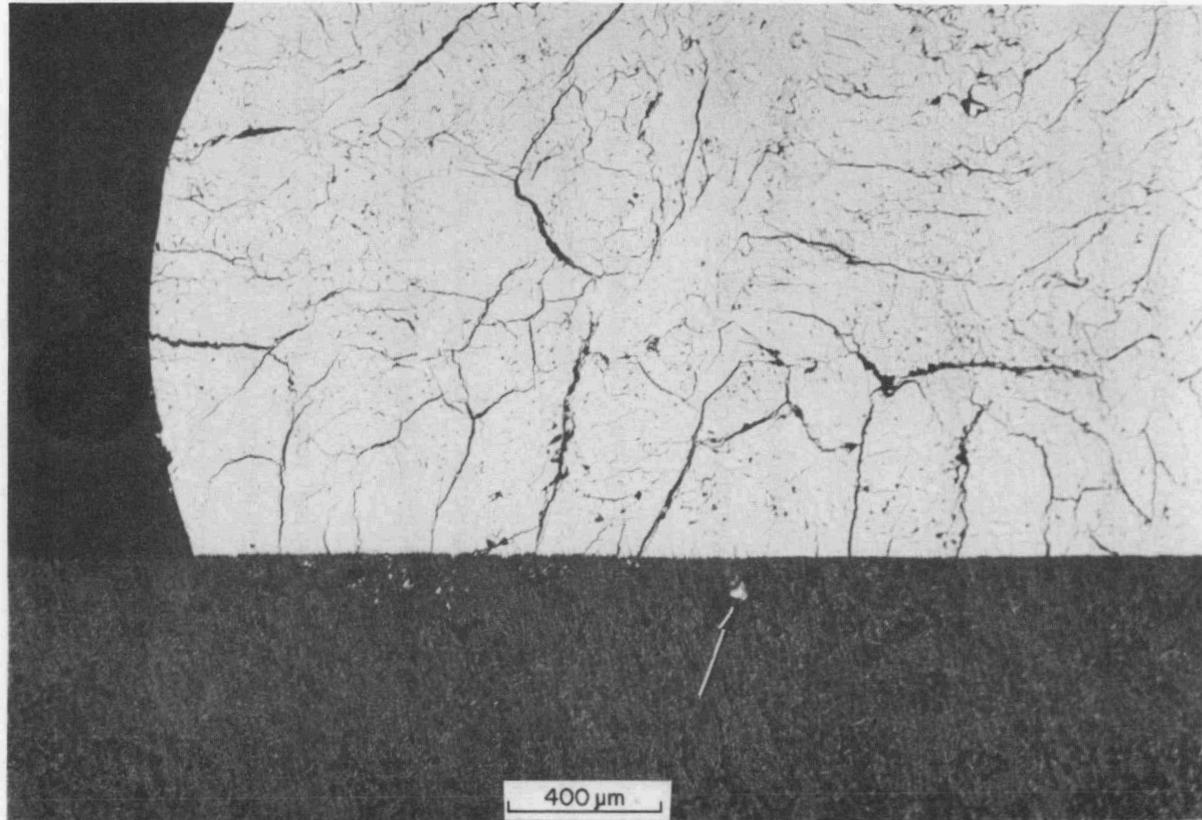


Fig. 11. Sessile drop of NiSi-NiSi₂ eutectic on spectroscopic carbon planchette. Note penetration of the metal into the graphite substrate (arrow). Wetting angle $\theta \approx 110^\circ$.

fractured into many pieces on cooldown to ambient. If that fracture had occurred during the run, on later meltings the molten NiSi-NiSi₂ would have run all over the furnace cavity.)

Although some penetration was seen in the sessile-drop test, it is clear that NiSi-NiSi₂ is not nearly as reactive toward graphite as is Mg₂Si-Si. On the other hand, the thermal expansion mismatch seems to be greater for this alloy. It is not known at this time whether the CTE of graphite could be reduced enough through microstructural control to make this system physically compatible. Preliminary tests to evaluate infiltration as a means of accommodating this CTE mismatch are reported in the following section.

3. ADVANCED CONCEPTS

3.1 TERNARY EUTECTIC ALLOYS

We have identified two ternary alloy systems with potentially interesting properties for advanced thermal energy storage systems, namely, Mg-Si-Ge and Ni-Si-Ge. These systems are attractive for the following reasons:

1. Germanium is very benign toward graphite. It is possible that the ternary alloys might combine the attractive thermal properties of the Mg-Si or Ni-Si systems with the desirable chemical characteristics of Ge.
2. Ternary eutectic systems provide a series of compositions whose individual melting temperatures can be chosen over a wide range, allowing greater latitude in designing the receiver and heat engine.
3. From a design standpoint, the thermal receiver might be more efficient if thermal energy storage modules with slightly different melting temperatures were distributed to eliminate dead spots that never melt or never freeze.
4. Because of the chemical similarity of Ge and Si, the Mg-Ge and Ni-Ge phase diagrams^{10,11} are similar to the Mg-Si and Ni-Si systems. Thus, the general features of the corresponding ternary systems can be predicted with some confidence.

3.1.1 The Mg-Si-Ge System

The Mg-Ge system¹⁰ contains the line compound Mg₂Ge, which melts at 1388 K. The Mg₂Ge-Ge eutectic occurs at 953 K and 82.3 wt % Ge. Note that Mg₂Ge is isostructural with Mg₂Si; its lattice parameter (6.390 Å) is very close to that of Mg₂Si (6.351 Å), suggesting that the line compound probably exists for all Ge-Si ratios. The Mg-rich side of

the Mg-Ge binary is virtually identical to that of the Mg-Si system. Based on these similarities, the probable Mg-Si-Ge ternary phase diagram is shown in Fig. 12.

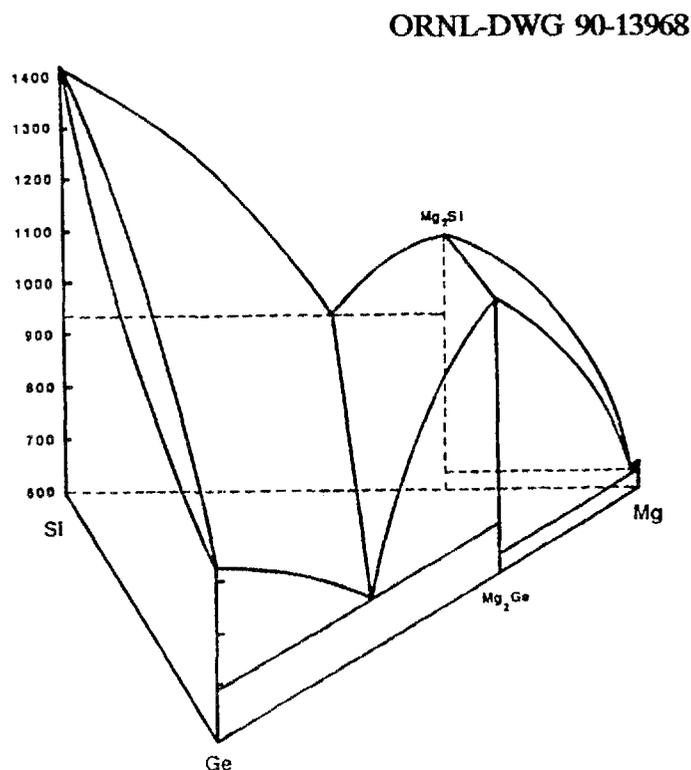
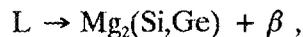


Fig. 12. Postulated Mg-Si-Ge ternary phase diagram, constructed from the respective binaries.

Compositions lying along the eutectic "trough" should have very narrow freezing ranges because they would solidify according to the eutectic reaction



where β is a Ge-Si solid solution.

As the first step toward characterizing the ternary system, Ge chips, Si chips, and Mg_2Si powder were mixed in a test capsule and melted at 1373 K. The composition was intentionally Ge-rich (Ge:Si = 4:1) and Mg-poor so that the first solid to appear on freezing should be Ge-Si solid solution. The liquid composition would then follow the eutectic trough shown in Fig. 12 and solidification would end with a reaction close to



The microstructure of this alloy (Fig. 13), along with qualitative energy dispersive X-ray analysis (EDX), confirms the general features of this alloy system. Proeutectic dendrites have a Si-rich core surrounded by more Ge-rich solid solution, as one would expect if freezing is too fast to allow homogenization in the solid phase. The interdendritic phase contains virtually no Si, as expected. (Note that EDX analysis is complicated by overlap of Ge $L\alpha$ and Mg $K\alpha$ lines. One can surmise the presence of Mg qualitatively by noting the relative intensities of Ge $L\alpha$ and $K\alpha$ lines.)

Further characterization of this alloy system should yield eutectic alloys with any desired melting point from 953 to 1213 K and large heats of fusion.

3.1.2 The Ni-Si-Ge System

Because the Ni-Si and Ni-Ge binary systems are so complex, it is not possible to construct the ternary diagram *a priori*. However, consideration of the binaries^{9,11} suggests that a range of compositions running approximately from the NiSi-NiSi₂ eutectic to the

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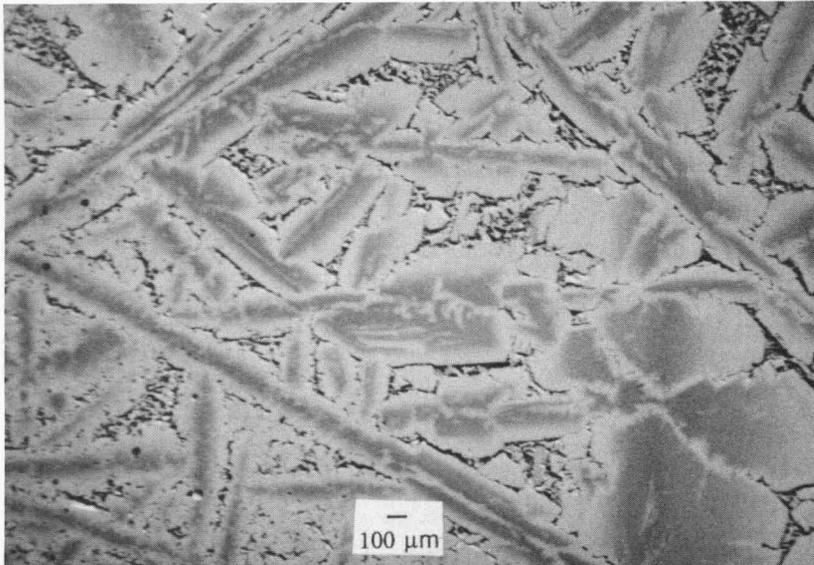


Fig. 13. Mg-Si-Ge alloy melted at 1373 K and furnace cooled. Microstructure consists of cored dendrites of Ge-Si solid solution surrounded by a Si-poor eutectic approximated by $Mg_2Ge + Ge$.

NiGe-Ge eutectic should yield relatively narrow melting points ranging from 1035 to 1239 K. Actual predictions are complicated by the fact that the compound NiSi₂ has no analog in the Ni-Ge system. Future work will include experimental characterization of the Ni-poor region of this ternary system, including thermal and microstructural analysis.

3.2 INFILTRATION TO CONTROL MELT GEOMETRY

It is well recognized that the long-term reliability of any PCM system depends on the ability to control the location of solid, liquid, and void within the individual PCM canisters. In the case of fluoride salts, the large volume change on melting implies a large void volume in the solid, which could lead to local hot spots during melting. Cotton and Sedgwick¹² describe the use of metal felt, which not only controls void distribution but also enhances thermal conductivity.

For our metallic PCMs, infiltration, or "wicking," has two main purposes. First, by controlling void distribution, one can avoid trapping liquid, which would catastrophically expand on freezing. Second, a carefully shaped wick could possibly maintain an appropriate gap to accommodate differential thermal expansion between the solid PCM and the graphite container. Another possible benefit (speculative at this time) is the distribution of nucleation sites throughout the PCM to help it freeze uniformly and to eliminate undercooling.

The wicking concept was tested on Mg₂Si-Si and NiSi-NiSi₂ eutectics because these alloys consistently broke their graphite containers (presumably from CTE mismatch) and both wet graphite readily. Two porous media were used as the wicks: CBCF insulation and reticulated carbon foam.

The CBCF material¹³ is made from chopped rayon fiber, which is carbonized and bonded with phenolic resin particles. CBCF shapes are made by vacuum molding from a water slurry, followed by drying and curing. Final carbonization of the resin is done at 1873 K. The resulting material has a nominal density of 0.2 g/cm³. It can be molded into simple shapes or into large plates from which smaller pieces can be cut with a sharp knife. Figure 14 shows the microstructure of a small CBCF preform before exposure to molten PCM.

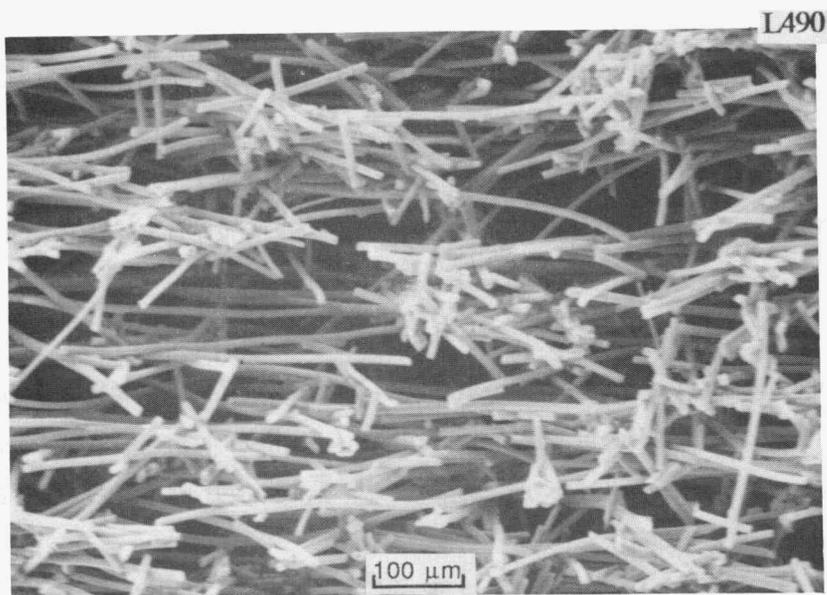
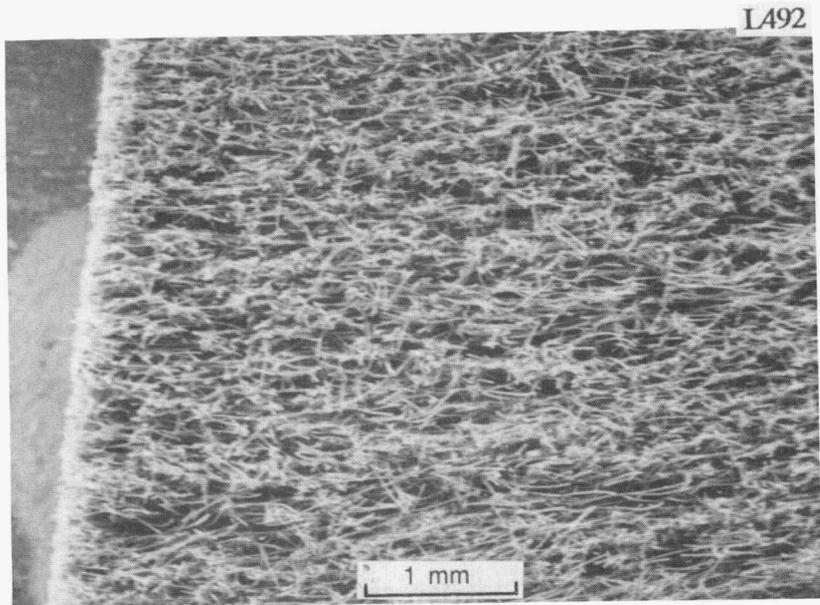


Fig. 14. Scanning electron micrographs of low-density CBCF composite. Top: general structure at 20× (bar = 1 mm). Bottom: detail showing fiber arrangement at 100× (bar = 100 μm).

CBCF was infiltrated with Mg_2Si -Si eutectic by placing a thin slab vertically in the center of a Grade 2 Stackpole 2020 test capsule and surrounding it with Mg_2Si powder and Si chips. After three cycles, although the charge had not fully settled and alloyed together, the eutectic that formed was drawn up into the CBCF. All expansion on freezing was then directed into the open space on either side of the wick (Fig. 15). This capsule is intact after 300 normal cycles as well as one intentional excursion to 1473 K.

Several larger preforms (cylinders and rectangular prisms) were cut from a CBCF plate; these were just small enough to fit inside a test capsule. These preforms were placed in a zirconia crucible, surrounded by Ni and Si (both powder and chips), and melted in vacuum at 1473 K. Samples held for 2 h at 1473 K were partially infiltrated by the NiSi-NiSi₂ eutectic. Samples held 3 h at 1573 K appeared to be completely infiltrated (Fig. 16). The infiltrated wicks are ready to be placed into test capsules for thermal cycling. Results of these tests will be reported later.

Carbon foam (Fig. 17) is made from a mixture of phenyl-furfuryl alcohol and urethane resin. These components are mixed in water with curing and blowing agents and cast into foam. The urethane resin is sufficiently viscous to support the foam structure, serving as a vehicle for the phenyl-furfuryl component. Heating to 1273 to 1573 K in argon or nitrogen burns off the urethane and carbonizes the phenyl-furfuryl alcohol. The resulting glassy carbon foam is graphitized at 2273 K.

The graphitized foam has a bulk density of 0.2 to 0.3 g/cm³ (about 10% theoretical density). There is a high degree of interconnectivity between the individual pores, although not every pore is directly accessible to each of its neighbors.

As the first step toward full infiltration tests, the ability of NiSi-NiSi₂ to wet carbon foam was evaluated by placing chunks (~1 cm) of the alloy on top of a 3 × 4 × 1 cm thick carbon-foam block. After several 2-h meltings at 1573 K in vacuum, the block became progressively more silvery. Scanning electron microscopy (Fig. 18) shows that the alloy penetrated throughout the pore spaces by wetting the surface of the carbon foam. Work will continue on this promising system in order to make samples with a much higher PCM loading for thermal-cycle testing.

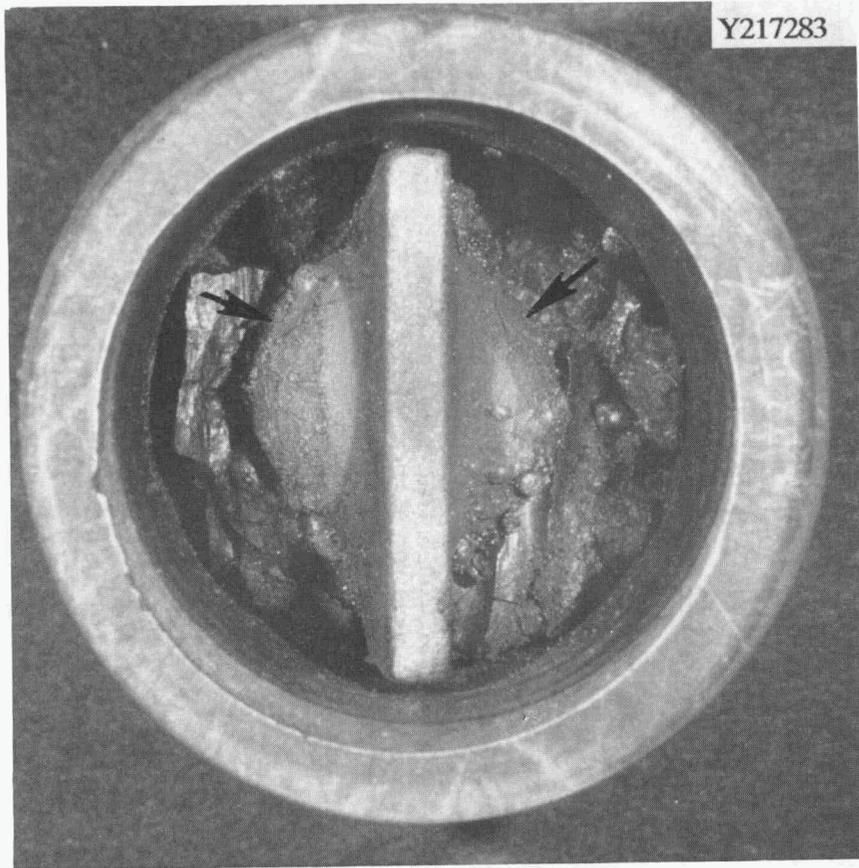


Fig. 15. Test capsule containing Mg_2Si -Si eutectic composition and a thin CBCF wick. Note how surface tension has drawn the molten alloy up into the wick (arrow).

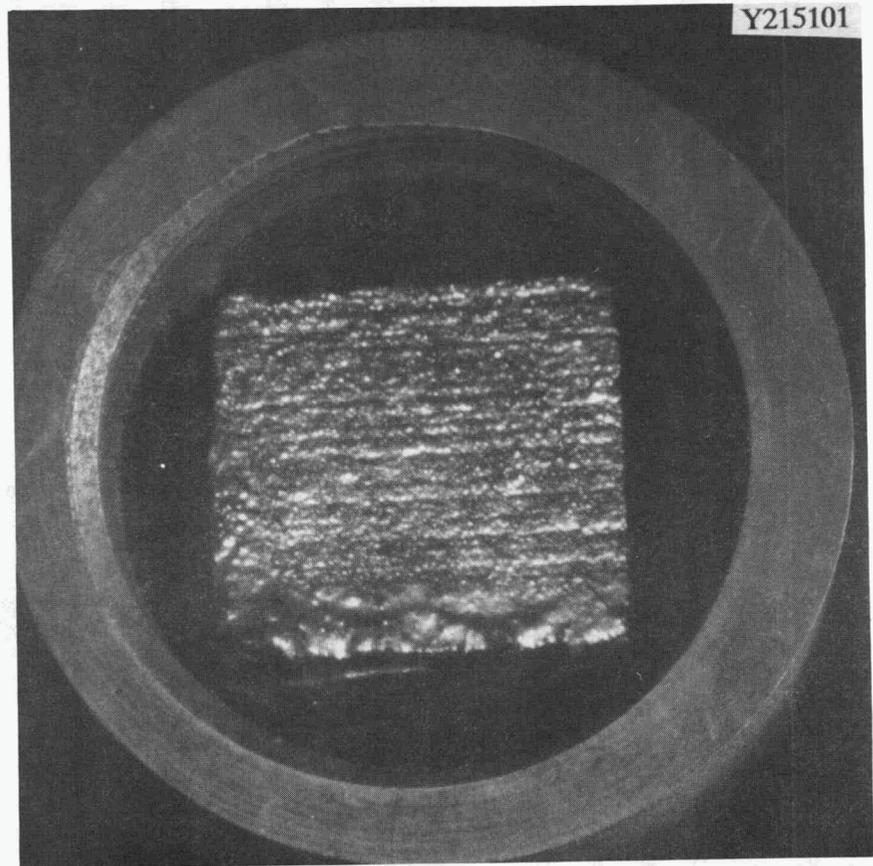


Fig. 16. Test capsule containing a rectangular CBCF preform infiltrated with $NiSi$ - $NiSi_2$ eutectic formed by melting the elements together at 1573 K for 3 h.

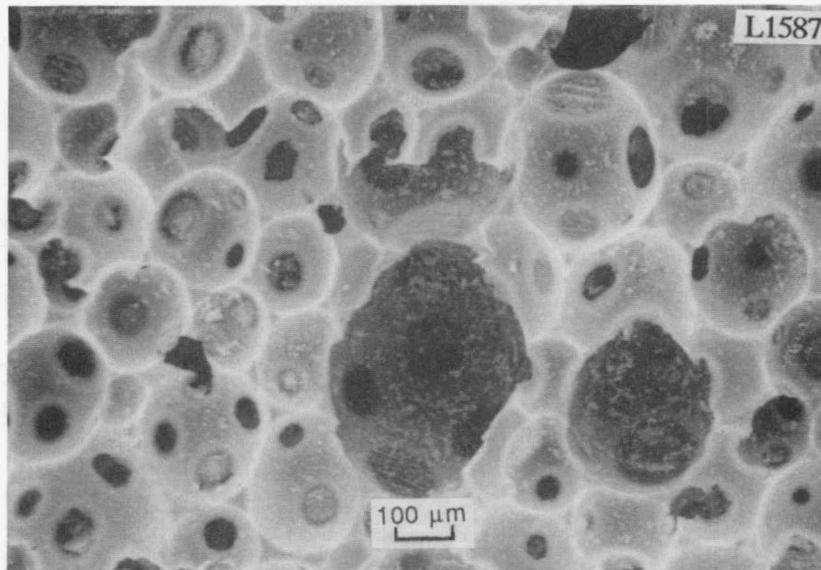
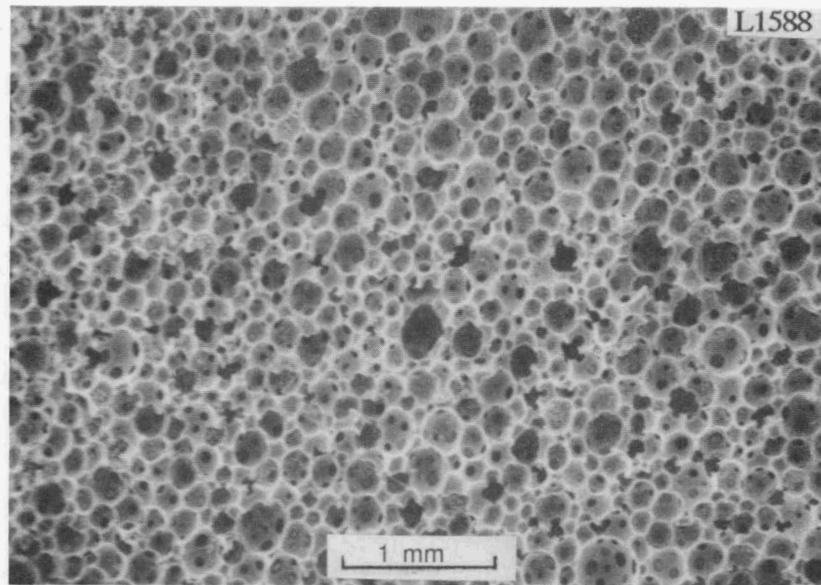


Fig. 17. Scanning electron micrographs of carbon foam. Top: overview of structure at 19 \times showing distribution of void diameters (bar = 1 mm). Bottom: detail showing open passages ($\sim 50 \mu\text{m}$) between voids at 80 \times (bar = 100 μm).

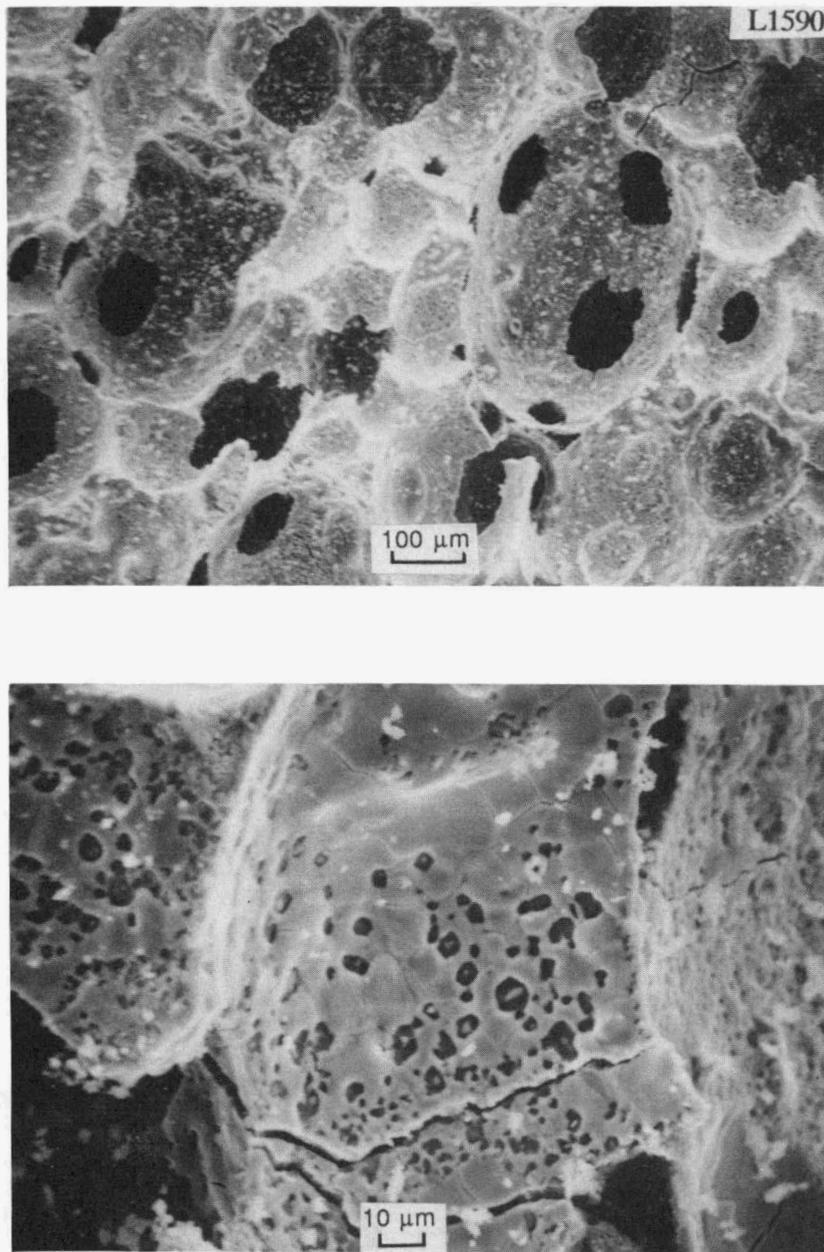


Fig. 18. Scanning electron micrographs of carbon foam exposed to NiSi-NiSi₂. One can see that the alloy has penetrated into the porous matrix by wetting the surface. Top: overview showing uniformity of the coating at 80× (bar = 100 μm). Bottom: detail at 500× showing that the alloy forms a nearly continuous film (bar = 10 μm).

4. CONCLUSIONS

1. A new thermal storage system using germanium metal as the PCM and POCO AXF-5Q graphite as the container shows potentially improved performance over baseline salt systems. Important features of the metal PCM are high thermal conductivity and high volumetric heat of fusion. The graphite container is easily fabricated, inexpensive, lightweight, and tolerant of much higher temperatures than Haynes 188.

2. Adding silicon to the germanium improves wetting of the graphite. Titanium does not promote wetting over the long term because of TiC formation.

3. Both NiSi-NiSi₂ and Mg₂Si-Si eutectics cannot be contained directly in graphite because of chemical interactions at the interface and large differences in thermal expansion in the solid state.

4. An infiltration technique has been developed that shows promise for eliminating CTE-driven fracture, provided that the PCM is an alloy that wets carbon. Both CBCF and reticulated carbon foam appear to be suitable matrices. Long-term testing is needed to assess the stability of these composite PCMs.

5. Ternary Mg-Si-Ge and Ni-Si-Ge alloys have potentially attractive properties for PCMs. In particular, these systems contain narrow-melting compositions that can be chosen over a wide range of possible melting points.

5. ACKNOWLEDGMENTS

R. L. Beatty first suggested the concept of a graphite capsule with large internal cavities to contain the PCM. J. R. Mayotte prepared metallographic specimens and photomicrographs. W. P. Eatherly provided much helpful advice on graphite chemistry and properties.

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6. REFERENCES

1. R. V. Boyle, M. G. Coombs, and C. T. Kudija, "Solar Dynamic Power Option for the Space Station," pp. 319-28 in *Proceedings of the 23rd Intersociety Energy Conversion Engineering Conference*, July 31-August 5, 1988, Denver, Colorado, ed. D. Y. Goswami, American Society of Mechanical Engineers, New York, 1988.
2. H. J. Strumpf, M. G. Coombs, and D. E. Lacy, "Advanced Space Solar Dynamic Receivers," pp. 357-65 in *Proceedings of the 23rd Intersociety Energy Conversion Engineering Conference*, July 31-August 5, 1988, Denver, Colorado, ed. D. Y. Goswami, American Society of Mechanical Engineers, New York, 1988.
3. E. J. Roschke, "Solar Dynamic Systems for Spacecraft Power Applications," AIAA Paper 80-0382, AIAA 24th Aerospace Sciences Meeting, Reno, Nev. (1986).
4. R. J. Lauf, T. B. Lindemer, and R. L. Pearson, "Out-of-Reactor Studies of Fission Product-Silicon Carbide Interactions in HTGR Fuel Particles," *J. Nucl. Mater.* 120(1), 6-30 (1984).
5. R. W. Olesinski and G. J. Abbaschian, "The Ge-C (Germanium-Carbon) System," *Bull. Alloy Phase Diagrams* 5(5), 484-86 (1984).
6. A. Adamson, *Physical Chemistry of Surfaces*, 4th Edition, pp. 332-68, Wiley, New York, 1982.
7. R. A. Crane and G. Bharadhwaj, "Thermal Performance of Specific Heat Receiver Concepts for Advanced Solar Dynamic Applications," pp. 943-51 in *Proceedings of the 24th Intersociety Energy Conversion Engineering Conference*, August 6-11, 1989, Washington D.C., ed. W. D. Jackson, Institute of Electrical and Electronics Engineers, New York, 1989.
8. M. Hansen, *Constitution of Binary Alloys*, pp. 916-17, McGraw-Hill Book Co., New York, 1958.
9. P. Nash and A. Nash, pp. 1755-56 in *Binary Alloy Phase Diagrams*, Vol. 2, T. B. Massalski, ed., Amer. Soc. for Metals, Metals Park, Ohio, 1986.
10. A. A. Nayeb-Hashemi et al., pp. 1222-23 in *Binary Alloy Phase Diagrams*, Vol. 2, T. B. Massalski, ed., Amer. Soc. for Metals, Metals Park, Ohio, 1986.

11. A. Nash and P. Nash, pp. 1230-32 in *Binary Alloy Phase Diagrams*, Vol. 2, T. B. Massalski, ed., Amer. Soc. for Metals, Metals Park, Ohio, 1986.
12. J. D. Cotton and L. M. Sedgwick, "Compatibility of Selected Superalloys with Molten LiF-CaF₂ Salt," pp. 917-21 in *Proceedings of the 24th Intersociety Energy Conversion Engineering Conference*, August 6-11, 1989, Washington D.C., ed. W. D. Jackson, Institute of Electrical and Electronics Engineers, New York, 1989.
13. G. C. Wei and J M Robbins, *Development and Characterization of Carbon-Bonded Carbon Fiber Insulation for Radioisotope Space Power Systems*, ORNL/TM-9381, Martin Marietta Energy Systems, Inc., Oak Ridge Natl. Lab., June 1985.

Appendix

Summary of Melting Tests

Test	PCM alloy	Capsule geometry	Material	Thermal cycles ^a
1	Ge	Sessile drop	Carbon planchette	1 h ^b
2	Ge-0.5% Ti	Sessile drop	Carbon planchette	4 h ^b
3	Ge-0.5% Si	Sessile drop	Carbon planchette	4 h ^b
4	NiSi-NiSi ₂	Sessile drop	Carbon planchette	4 h ^b
5	Ge	Test	2020 Grade 1	637+
6	Ge	Test	2020 Grade 2	544
7	Ge-0.5% Ti	Test	2020 Grade 1	637+
8	Ge-0.5% Ti	Test	2020 Grade 1	637+
9	Mg ₂ Si-Si	Test	2020 Grade 2	3
10	Mg ₂ Si-Si	Test	2020 Grade 1	7
11	NiSi-NiSi ₂	Test	2020 Grade 2	1
12	Ge-0.5% Si	Test	2020 Grade 2	52
13	Ge-0.5% Ti	Test	2020 Grade 2	52
14	Ge	Prototype, drilled	AXF-5Q	3
15	Ge	Prototype, annulus	2020 Grade 1	482+
16	Empty	Test	2020 Grade 1	591+
17	Ge	Prototype, drilled	AXF-5Q	3
18	Ge	Prototype, annulus	AXF-5Q	495+
19	Mg ₂ Si-Si	Test	2020 Grade 1 w/CBCF "wick"	308+

^aEach cycle includes 1 h at 773 K and 2 h at 1273 K, as described in the test. "+" indicates that the capsule was intact after the indicated cycle and tests are continuing.

^bSessile-drop tests were held at 1273 K for the time indicated (no cycling).

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